第5講義
表面利用

活性炭の応用

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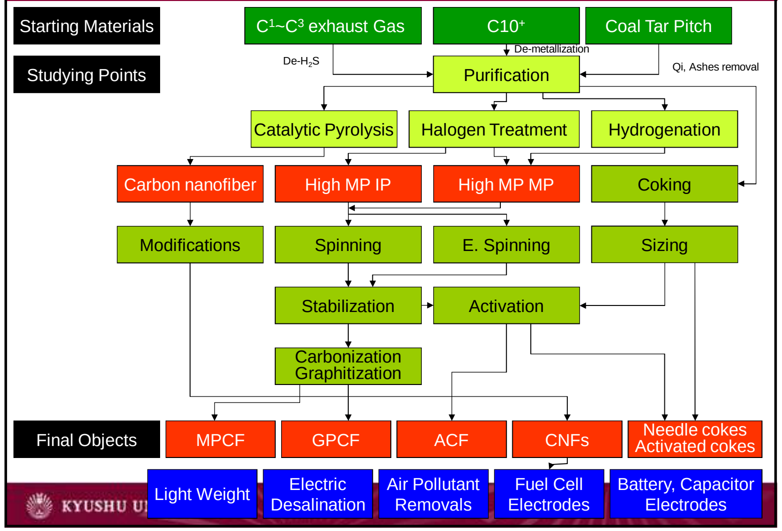
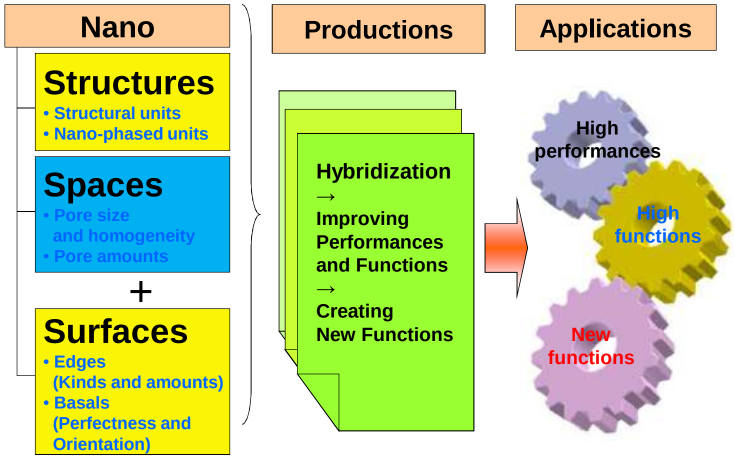
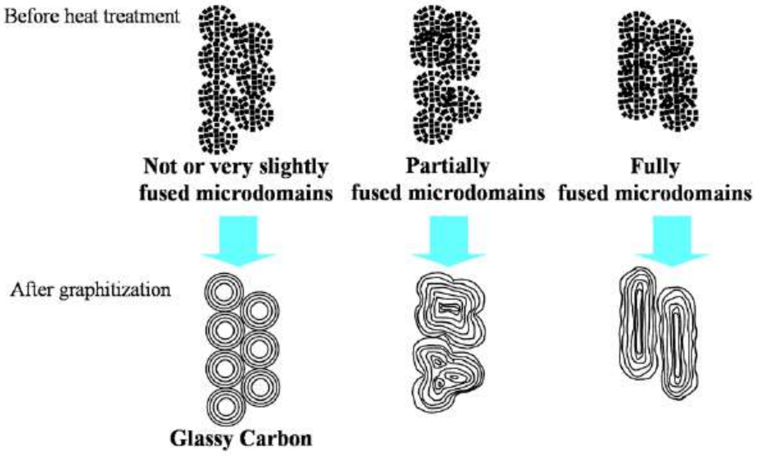
More concrete connections with raw materials to FCs



Understanding carbon structures：Carbon nano-world



Basic structure and structural control of carbon



IAMS, Kyushu University "Axial nano-scale microstructure in the graphitized fiber inherited from liquid crystal mesophase pitch" *Carbon*, 34, 83-88 (1996) S. H. Yoon, Y. Korai, K.Yokogawa, S. Fukuyama, M. Yoshimura, I. Mochida

Nanoscopic Structure of PAN Based CF

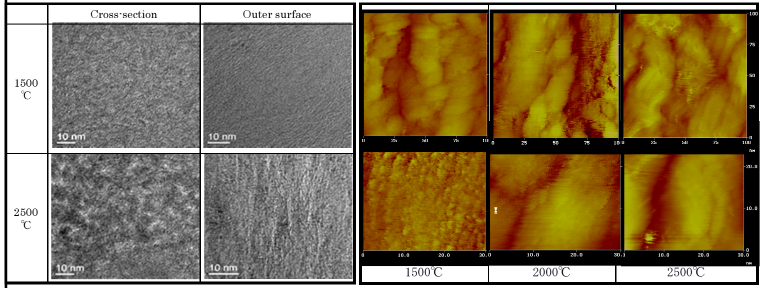
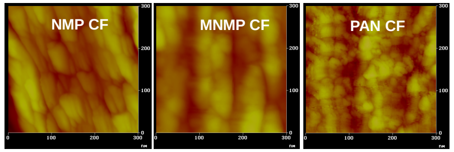
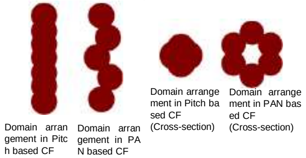


Figure SEM & STM images of heat treated PAN based CFs at 1500、2000、2500℃





“Structural comparison of mesophase and PAN based carbon fibers”
S.H. Hong, S. H. Yoon, I. Mochida *Carbon2006*、 (2006、7) England

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Key Materials for Energy and Environmental Devices

Carbon is an Indispensable Material for Energy related Devices

Best Structure for Best Performance
↕
Best Selection

Best Selection

Scientific Cycle

- Structural Understanding
- Structure Preparation
- Working Mechanism
  - Molecular Level
  - Electrochemical
  - Catalytic / Kinetics
  - Molecular / Heat Transfer

Carbon

Preparation of Activated Carbons

Selection of Precursor

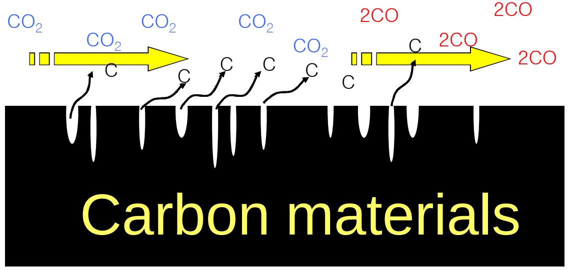
- *Pore Framework / Density*
- *Properties of Pore Wall, Composition / Graphitic Extent*
- *Reactivity at Activation*
- *Non-graphitizable precursors like polymer, biomass and isotropic coke for usual AC or ACF*
- *Graphitizable precursors like anisotropic cokes or mesophase pitch for EDLC electrode materials*

Activation Procedures

- *CO_2, H_2O*
- *Alkali Hydroxides / Carbonates; More Research*
- *Selective Catalytic Gasification ; Catalyst Control*

❖Very Large Surface Area > 3000 m^2/g
❖Adequate Pore And Wall

Activation (Making small pores in the carbon materials)



Activation reagents

- Air, CO2, Steam
- KOH (NaOH), ZnCl2

KOH Activation

Catalytic progress
Under K2O

$2KOH \rightarrow K_2O+H_2O$
(Dehydration)

$C+H_2O \rightarrow H_2+CO$
(Water gas reaction)

$CO+H_2O \rightarrow H_2+CO_2$
(Water gas shift reaction)

$K_2O+CO_2 \rightarrow K_2CO_3$

Carbon is mainly consumed as a form of K_2CO_3

$K_2O+H_2 \rightarrow 2K+H_2O$

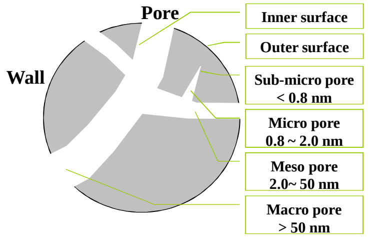
$K_2O+C \rightarrow 2K+CO$

K compounds are reduces as a metal

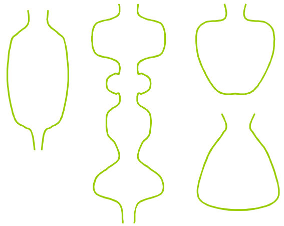
- Higher surface area compared to the steam activation
- Almost no productions of CO and CO_2
- K metal intercalation: higher diffusivity than steam molecule

KOH	K_2O	K_2CO_3	K
MP：380℃ BP：1324℃	MP：490℃（350℃, KO and K ）	MP：891℃	MP：64℃ BP：774℃

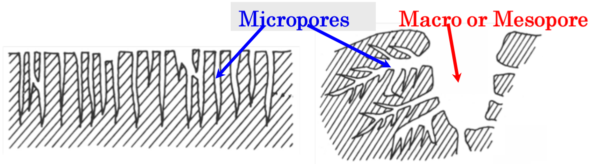
Activated carbon : Images of pores



Classification of surface and pores

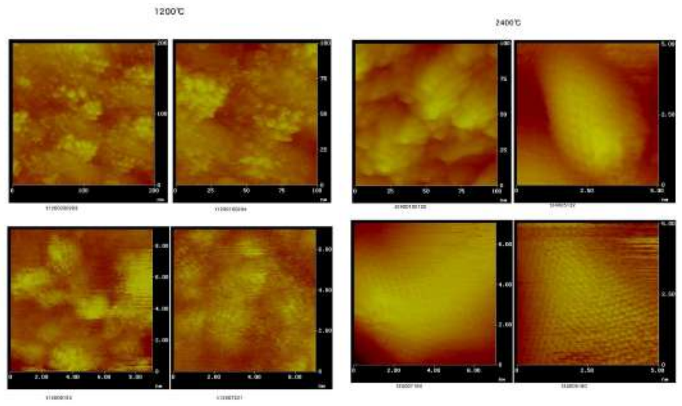


Schematic shapes of pores



Schematic pore images of activated carbon fiber and activated carbon

Glassy carbons from Novolak resin



Structure of Activated Carbon

Surface Area, Pore：*Depth & Volume*
Surface Structure
Surface Chemistry
Based and Edge Plane, Substituents
Hetero Atoms in Hexagon

Carbon Structure of Wall
Micro, Nano, Macro Structure of Carbon Wall
-Graphitization Extent
-Domain Structure
Density, Reactivity (Activated Surface)
→ *Precursor : Structure and Reactivity*

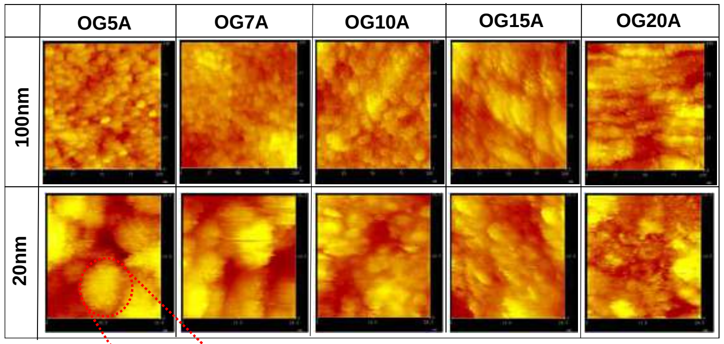
Some properties of pitch based ACFs

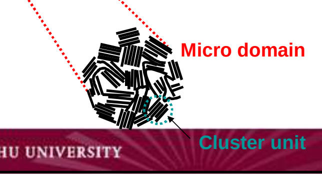
♦ Pitch-based activated carbon fibers (ACFs) used in this study.
OG5A, OG7A, OG10A, OG15A and OG20A were provided by Osaka Gas Co.

Some physical properties of ACFs given by Osaka Gas Co.

	OG5A	OG7A	OG10A	OG15A	OG20A
Specific surface area (m²/g)	480	850	1300	1725	2000
Average pore size (nm)	1.4	1.6	1.8	1.9	2.2
Average pore volume (cm³/g)	–	0.38	0.54	0.81	1.13

STM images of ACFs



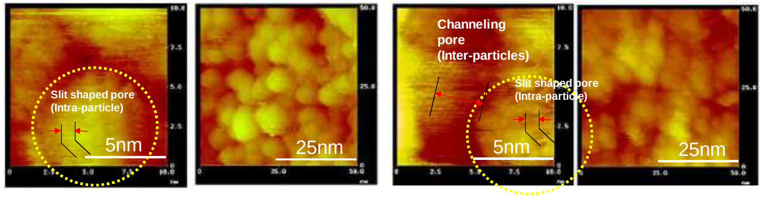


♦ACFs consist of structural units of micro domain with diameter of around 5nm.

STM images of ACFs

In order to remove oxygen containing functional groups for removing the heterogeneous effect of STM, OG7A and OG20A were heat-treated at 800^{O}C in a hydrogen atmosphere (H_2 / He =1/4).

OG7A-800H OG20A-800H



♣**Vacant spaces between the two domains of OG20A are larger than that of OG7A.**

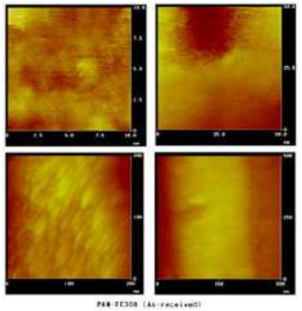
♣**Domain size of OG20A is a little smaller than that of OG7A.**

♣**Slit type pores were observed in domains of OG7A and OG20A.**

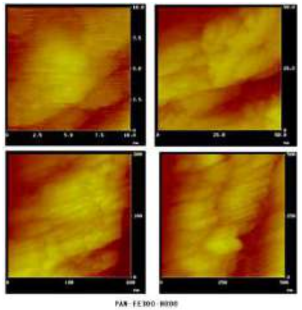
♣**It can be presumed that almost pores larger than 2nm nucleated by the inter-particle mechanism.**

STM images of PAN Based ACFs

FE300 As received



FE300 -800H



♣ **FE300 is also composed of micro-domain structures.**
♣ **Basic structural units of PAN based ACF are more difficult to observe because of high heterogeneous effects of surface heterogeneous atoms and many defects.**

Activation of carbons

Precursor of ACF has been composed of nano-structural primary units

Structural factor should be considered for the better understanding of activated carbons and their applications.

Size and arrangement of BSU
Etching and diffusion of oxidative agent against BSU

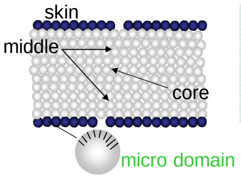
Pores from intraparticles (Slit shaped? Micropores less than 2 nm)

Pores from interparticles (Channeled shaped having wider pore size distributions (0.2 ~ 50 nm)

Mechanism of the pore formations in ACFs

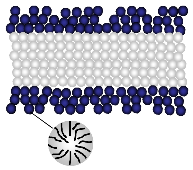
A model for cross section of ACFs

OG5A



♣Only skin is activated, homogeneous narrow pore exist on the surface of domains.
♣The activation does not reach to the middle and core parts.
↓
OG5A has smallest surface area but homogeneous pores.

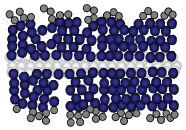
OG7A



♣Skin and middle parts are activated.
♣Pores in domain become wider and longer than that of OG5A because of the pores formed by inter-domain mechanism.
♣Pores are formed by the intra-domain and inter-domain mechanisms.
↓
Heterogeneous pores exist.

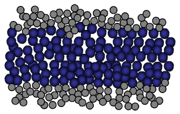
Mechanism of the pore formation in ACFs

OG10A



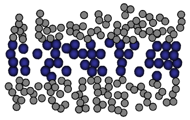
♣The activation proceeds to near the core.
♣Domains in the surface become smaller by activation.
↓
OG10A have various size pores.

OG15A



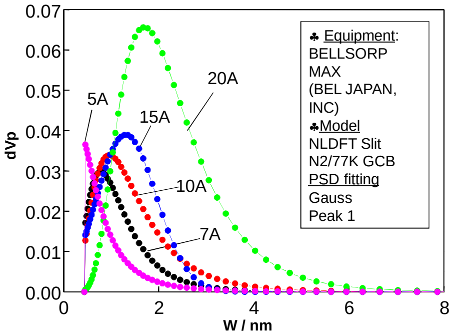
♣Whole parts of skin, middle and core are activated.
♣The pores from the intra-domain become larger by activation.
♣Owing to some domains burning, fiber diameter becomes smaller than OG5A, 7A and 10A.
↓
More homogeneous pore distribution than that of OG10A

OG20A



♣Over 70% of domains are burned-off.
♣Pore portions from interparticle nucleation becomes larger
↓
♣OG20A diameter become smaller than OG15A.
OG20A have largest surface area but heterogeneous pores.

Pore distribution using NLDFT method



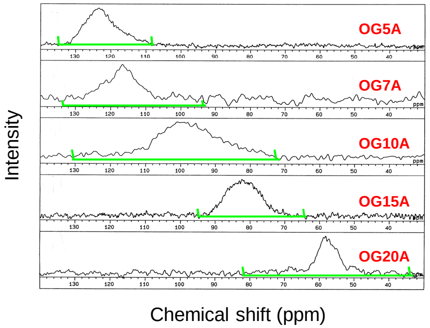
♣Advantage of DFT method
Determination of a pore size distribution in the wide range of pore size from micro pore to meso pore.

♣Disadvantage
This method needs
>measurement in low pressure
>assumption of pore structure
>homogeneity of adsorbent surface

	OG5A	OG7A	OG10A	OG15A	OG20A
pore size at peaks (nm)	Less than 0.44	0.72	0.98	1.28	1.86
Pore distribution range in this DFT profile (nm)	~2.8	~3.0	~5.0	~3.0	0.44~7.0

♦The pore size at peak increased in the order of OG5A<7A < 10A < 15A < 20A.

^{129}Xe-NMR spectrum of ACFs



	pore size at peak (nm)	Pore distribution width (nm)
OG5A	0.84	0.81~0.98
OG7A	0.89	0.79~1.15
OG10A	1.04	0.82~1.50
OG15A	1.23	1.15~1.55
OG20A	1.49	1.32~3.84

- ◆ Chemical shift at highest intensity, corresponding to pore size, increased in the order of OG5A<7A < 10A < 15A < 20A.
- ◆ 5A and 15A exhibited narrow peaks. →relatively homogeneous pore size
- ◆ Peak broadening can be seen for 7A, 10A and 20A. → heterogeneous pore size

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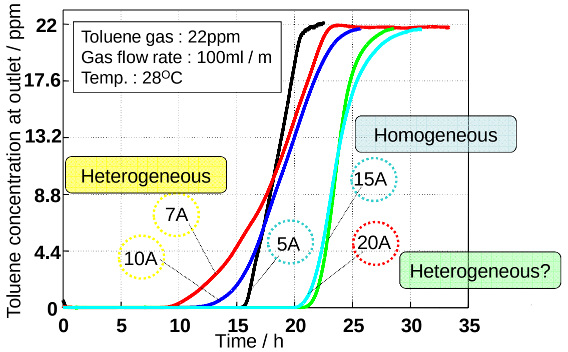
Toluene adsorption characteristics of ACFs



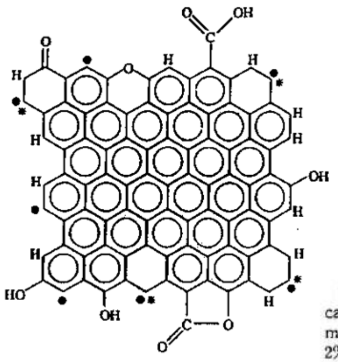
- ◆The slopes of breakthrough curves for 5A and 15A were steeper than those of curves for 7A and 10A.
- ◆ACFs with homogeneous pores (5A and 15A) showed rapid toluene adsorption and larger capacity per unit area and longer breakthrough time, whereas ACFs with heterogeneous pores (7A and 10A) showed slow toluene adsorption and smaller capacity per unit area.

Some Properties of ACFs

Pitch based ACF	BET (m^2 / g)	Elemental analysis (wt %) C	H	N	O	N / C
OG5A	563	92.4	0.6	0.7	6.0	0.007
OG7A	901	93.0	0.6	0.8	5.4	0.007
OG10A	1085	95.3	0.6	0.5	3.4	0.004
OG15A	1606	95.2	0.6	0.3	3.4	0.003
OG20A	1924	94.1	0.6	0.4	4.8	0.003

PAN based ACF	BET (m^2 / g)	Elemental analysis (wt %) C	H	N	O	N / C
FE100	450	70.9	2.0	8.4	17.3	0.102
FE200	650	72.5	1.8	4.8	17.9	0.057
FE300	880	74.3	1.6	3.3	17.2	0.038
FE400	1020	76.8	1.6	2.3	19.4	0.026

Surface Oxygen Functional Groups of ACF



This structure is representative of an activated carbon with a crystallite width of 15 Å and an elemental analysis (by weight) of 87.5% C, 11.3% O, 1.2% H, ● represents an unpaired σ electron ; ●* represents an "in-plane σ pair" with * being a localized π electron, (Radovie)

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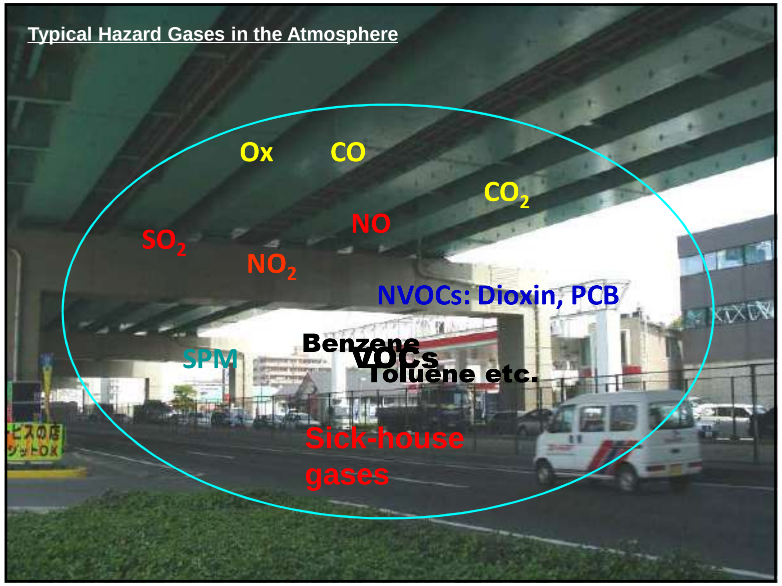
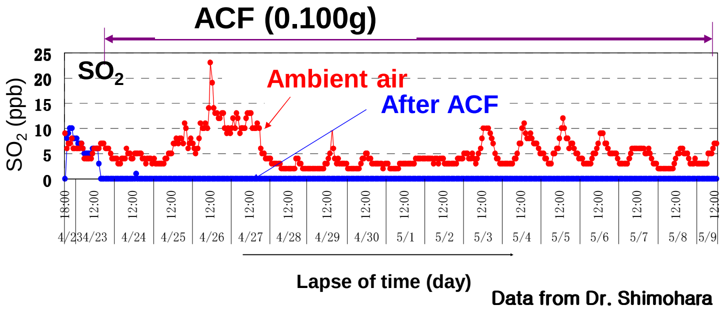
Removal of SOx and NOx Using ACFs

Typical Hazard Gases in the Atmosphere

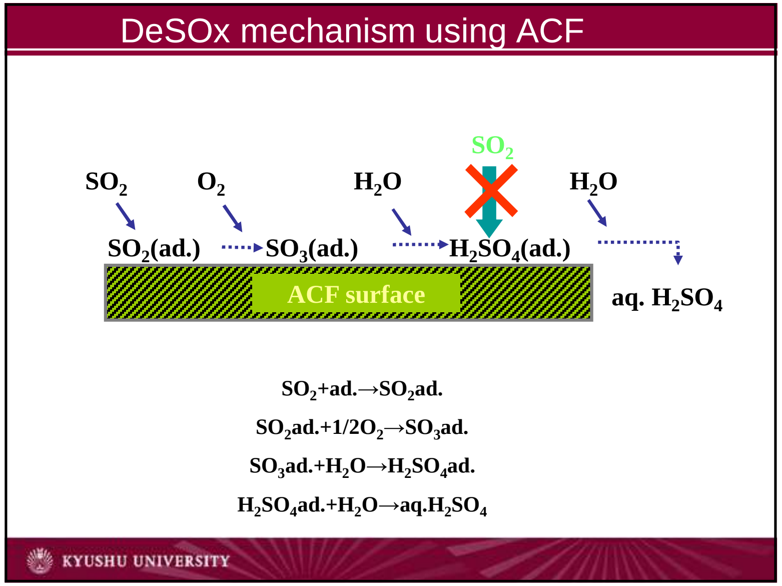


ACF

SO_2 ⇨ SO_4^{2-}



Data from Dr. Shimohara
Of Fukuoka H & E Institute



DeSOx mechanism using ACF

SO_2+ad.→SO_2ad.

SO_2ad.+1/2O_2→SO_3ad.

SO_3ad.+H_2O→H_2SO_4ad.

H_2SO_4ad.+H_2O→aq.H_2SO_4

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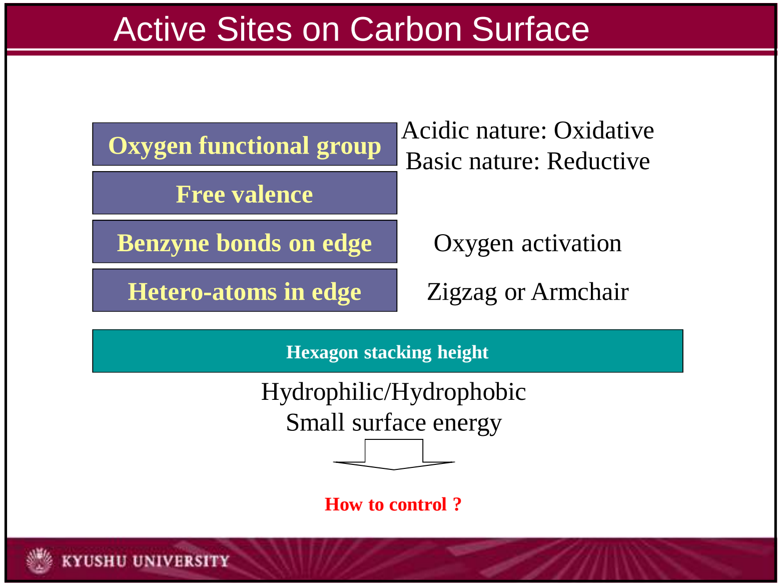

Active Sites on Carbon Surface

Oxygen functional group	Acidic nature: Oxidative Basic nature: Reductive
Free valence	
Benzyne bonds on edge	Oxygen activation
Hetero-atoms in edge	Zigzag or Armchair

Hexagon stacking height

Hydrophilic/Hydrophobic
Small surface energy

How to control ?

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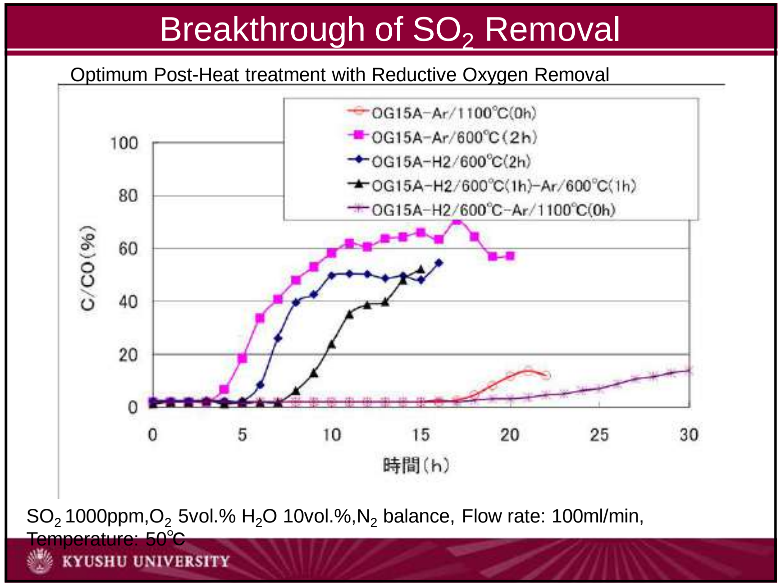

Breakthrough of SO_2 Removal

Optimum Post-Heat treatment with Reductive Oxygen Removal

SO_2 1000ppm,O_2 5vol.% H_2O 10vol.%,N_2 balance, Flow rate: 100ml/min, Temperature: 50℃

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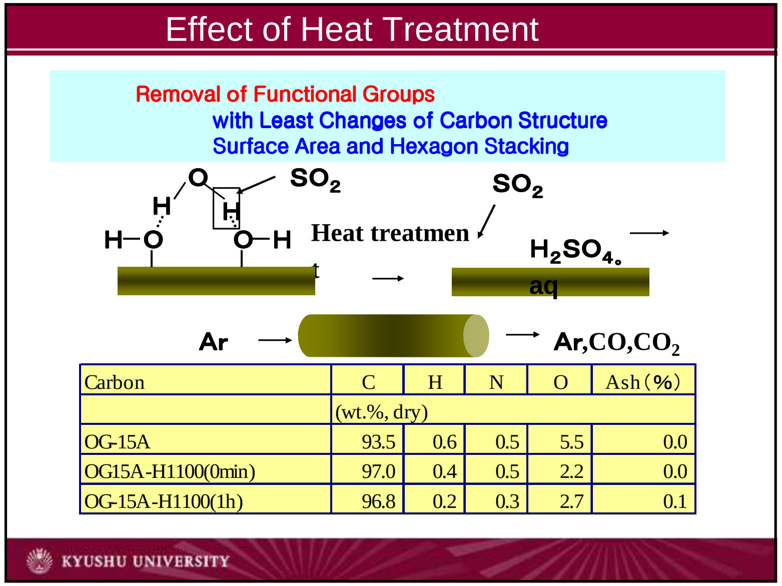

Effect of Heat Treatment

Removal of Functional Groups
with Least Changes of Carbon Structure
Surface Area and Hexagon Stacking

Carbon	C	H	N	O	Ash(%)
	(wt.%, dry)				
OG-15A	93.5	0.6	0.5	5.5	0.0
OG15A-H1100(0min)	97.0	0.4	0.5	2.2	0.0
OG-15A-H1100(1h)	96.8	0.2	0.3	2.7	0.1

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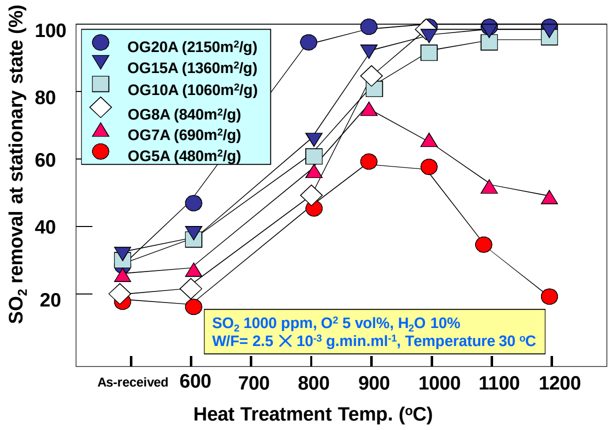
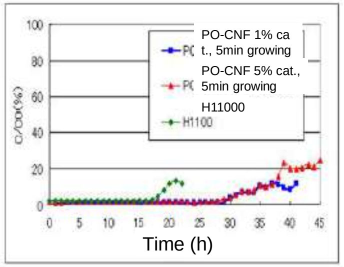
The Effect of Heat Treatment for SO_2 Removal



DeSOx by ACF and CNF-ACF Composite

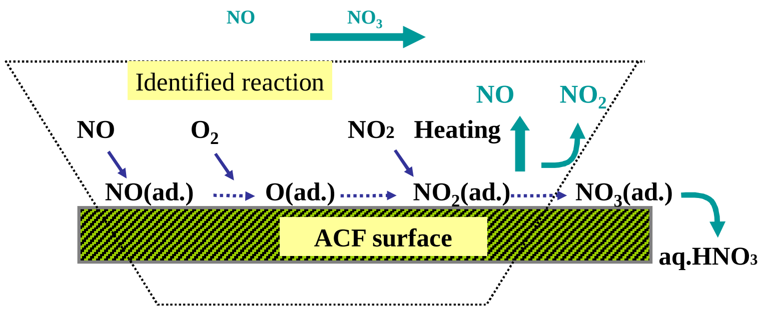
DeSOx Properties of ACF and ACF−CNF



DeSOx condition: SO_2 1000ppm, O_2 5vol%, H_2O 10vol%,
N_2 balance、Total flow rate: 100 ml/min
Reaction Temperature: 50 ℃

PDU for SOx Removal by ACF

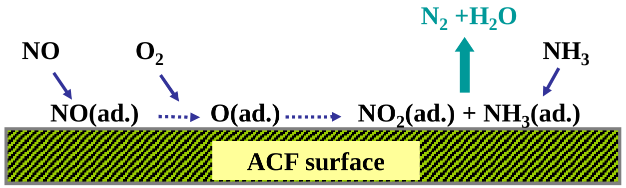
NO & NO_2 Oxidation over ACF



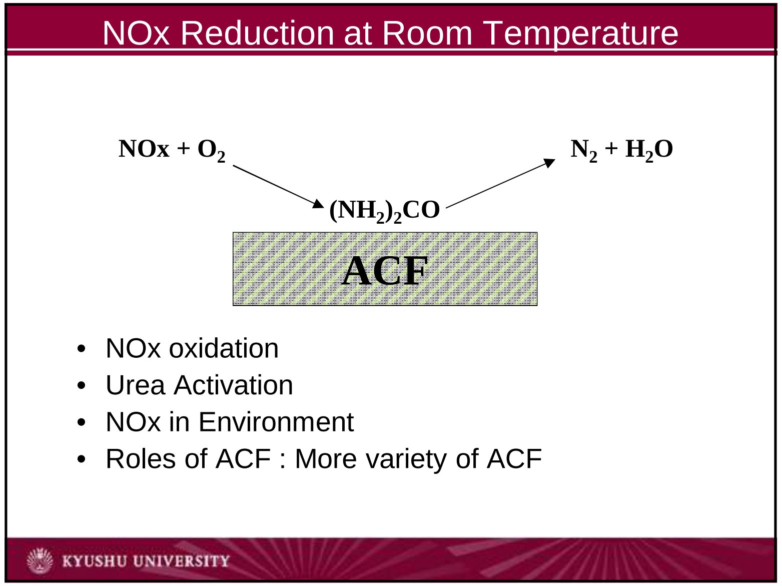
Strong Inhibition of H_2O

The oxidation of NO_2 always produces NO And NO_3^- through the disproportionation.

The Mechanism of NO Reductive Removal



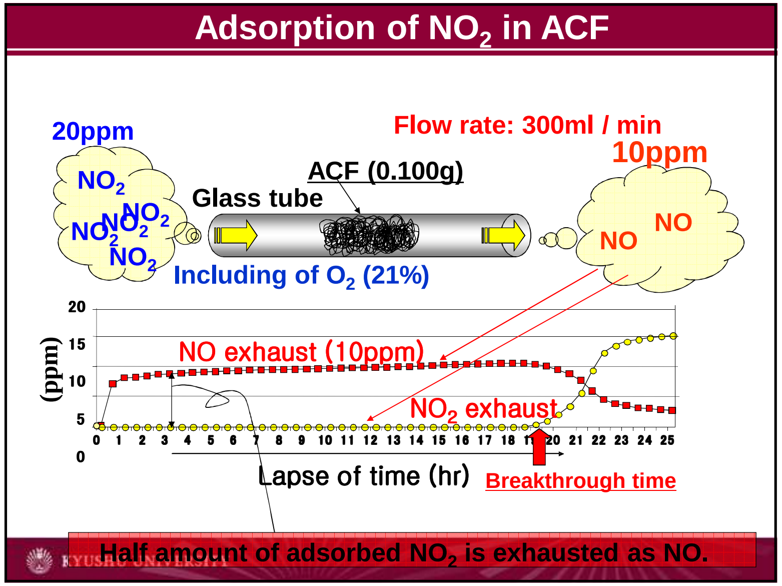
The mechanism of NO removal consists of adsorption and oxidation of NO into NO_2 which is reduced with NH_3



NOx Reduction at Room Temperature

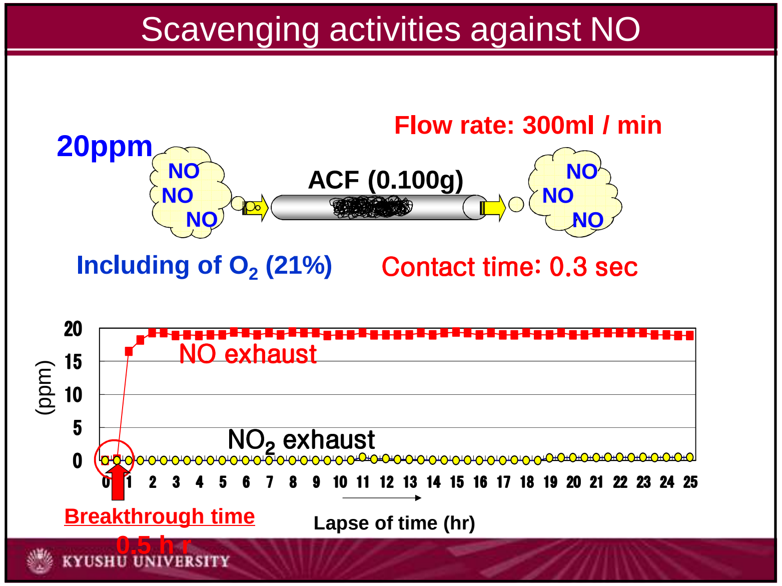
$NOx + O_2 \rightarrow (NH_2)_2CO \rightarrow N_2 + H_2O$

- NOx oxidation
- Urea Activation
- NOx in Environment
- Roles of ACF : More variety of ACF

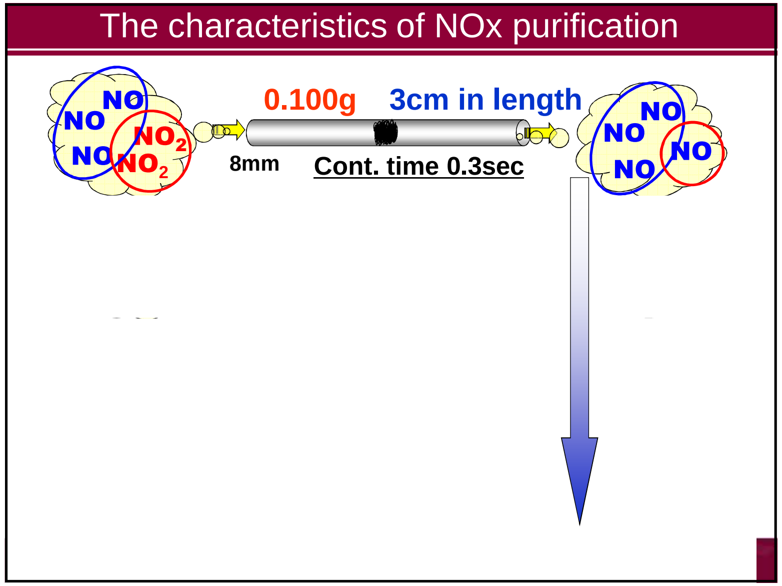


Adsorption of NO_2 in ACF

Half amount of adsorbed NO_2 is exhausted as NO.

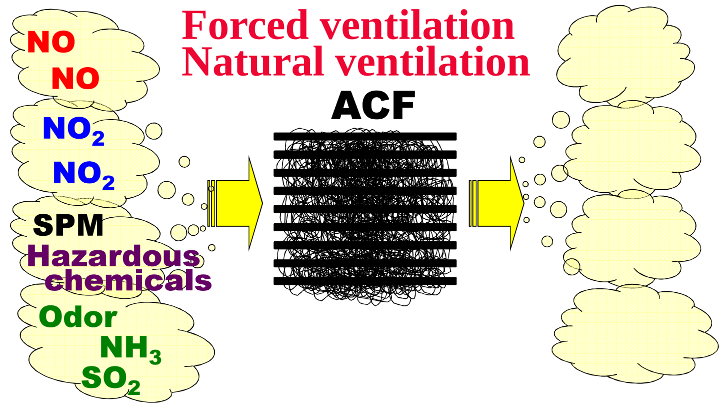


Scavenging activities against NO



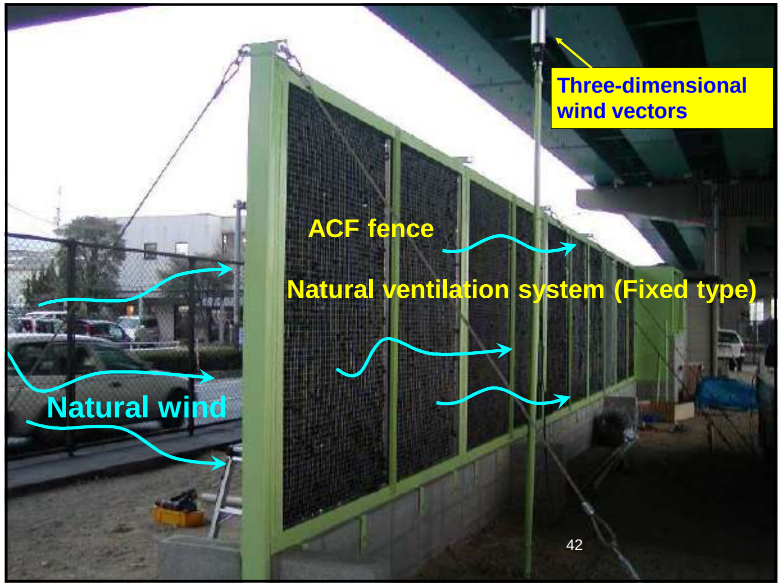
The characteristics of NOx purification

Characterization of ACF purification



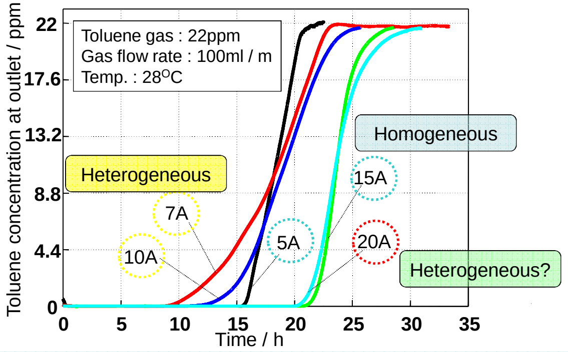
Forced ventilation
Natural ventilation

Room temperature, ozonizer is no need, no light irradiation, compact design



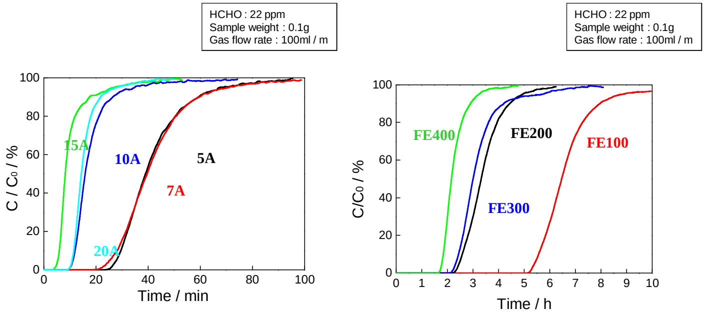
Removal of Sick-house gases
using ACF/MnO_2

Toluene adsorption characteristics of ACFs



- ♦The slopes of breakthrough curves for 5A and 15A were steeper than those of curves for 7A and 10A.
- ♦ACFs with homogeneous pores (5A and 15A) showed rapid toluene adsorption and larger capacity per unit area and longer breakthrough time, whereas ACFs with heterogeneous pores (7A and 10A) showed slow toluene adsorption and smaller capacity per unit area.

HCHO adsorption characteristics of ACFs

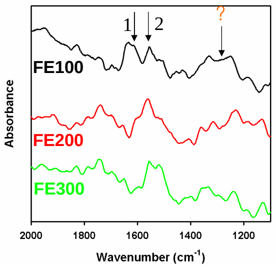


Break through time

◆Pitch-based ACF : 15A < 20A < 10A < 7A < 5A

◆PAN-base ACF : FE400 < FE300 < FE200 < FE100

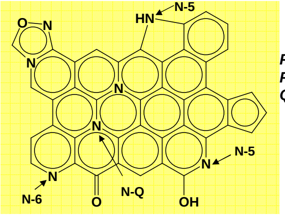
Micro ATR-FTIR spectrum of FE series



1) *Pyridinic nitrogen band*
2) *Internal standard*

	Micro-ATR FTIR analysis			EA and XPS analysis		
	I.S[a]	Pyridinic N[b]	RCP[c]	N / C ratio	Pyridinic N[b]	RCP[c]
FE100	279	134	0.48	0.118	0.737	0.087
FE200	276	108	0.39	0.066	1.050	0.069
FE300	332	70	0.21	0.044	0.734	0.032

[a] I.S : Internal standard (1560cm^{-1})
[b] Wavenumber related with pyridinic nitrogen : 1610~1600cm^{-1}
[c] RCP : relatives contents of pyridinic nitrogen



Pyridinic N (N-6)
Pyrrolic or Pyridinic N (N-5)
Quaternary N (N-Q)

Micro-ATR FTIR analysis

Relative amount of pyridinic nitrogen functional groups for PAN based ACFs by micro-ATR FTIR analysis

Sample	Internal Standard	Pyridinic N [a]	Internal Standard /Pyridinic N
FE100	279	134	0.48
FE200	276	108	0.39
FE300	332	70	0.21
FE400	330	64	0.19

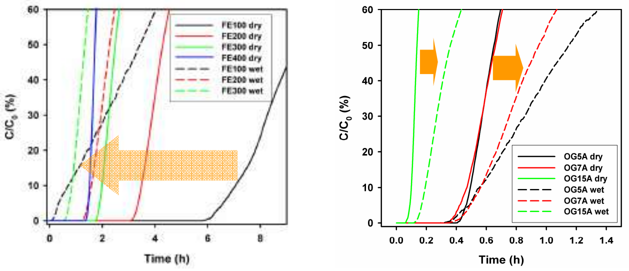
---

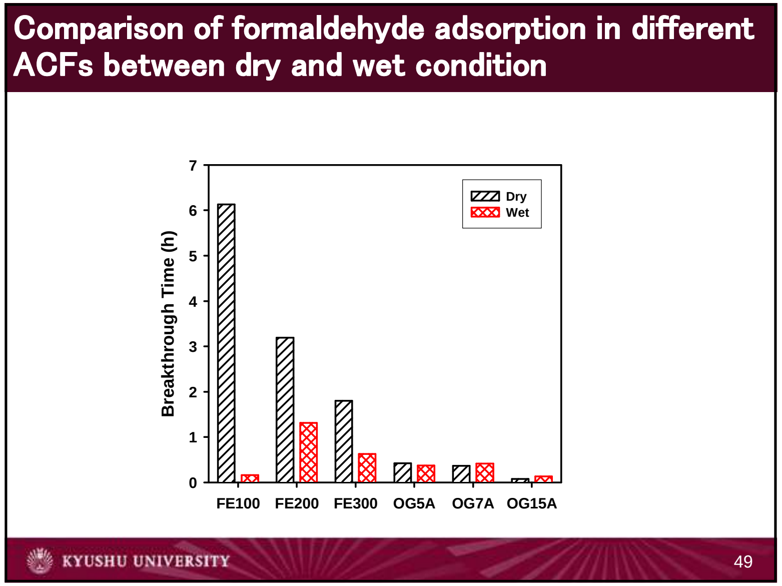
[a] Wavenumber related with pyridinic nitrogen: 1610 ~ 1600 cm^{-1}

Breakthrough curves of formaldehyde adsorption

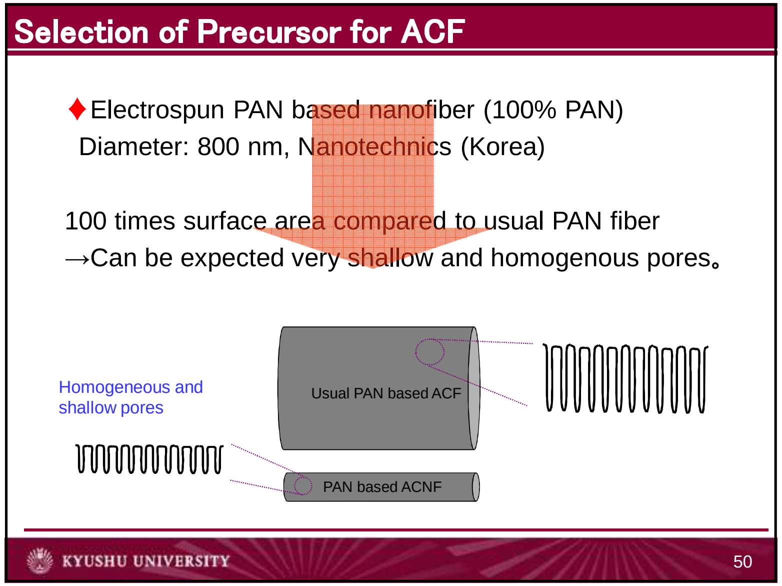
WATER Competitive adsorption decreases the adsorption amount of HCHO

Dry condition (solid line) and wet condition (dashed line) for the different kinds of a) FE series and b) OG series





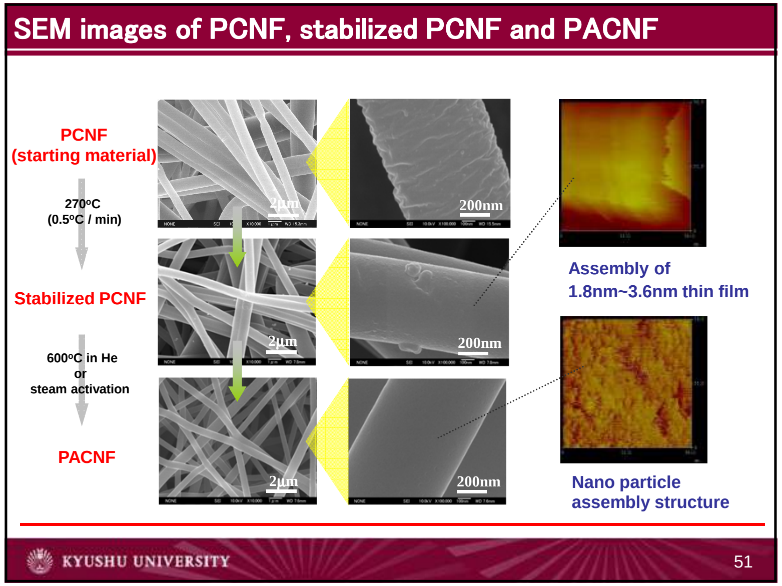
Comparison of formaldehyde adsorption in different ACFs between dry and wet condition



Selection of Precursor for ACF

♦Electrospun PAN based nanofiber (100% PAN)
Diameter: 800 nm, Nanotechnics (Korea)

100 times surface area compared to usual PAN fiber
→Can be expected very shallow and homogenous pores。



SEM images of PCNF, stabilized PCNF and PACNF

PCNF (starting material)

270°C
(0.5°C / min)

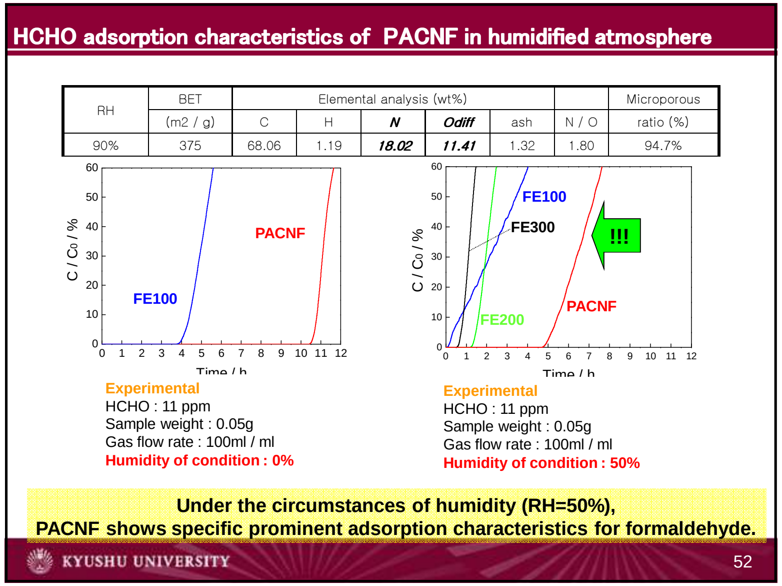
Stabilized PCNF

600°C in He
or
steam activation

PACNF

Assembly of 1.8nm~3.6nm thin film

Nano particle assembly structure



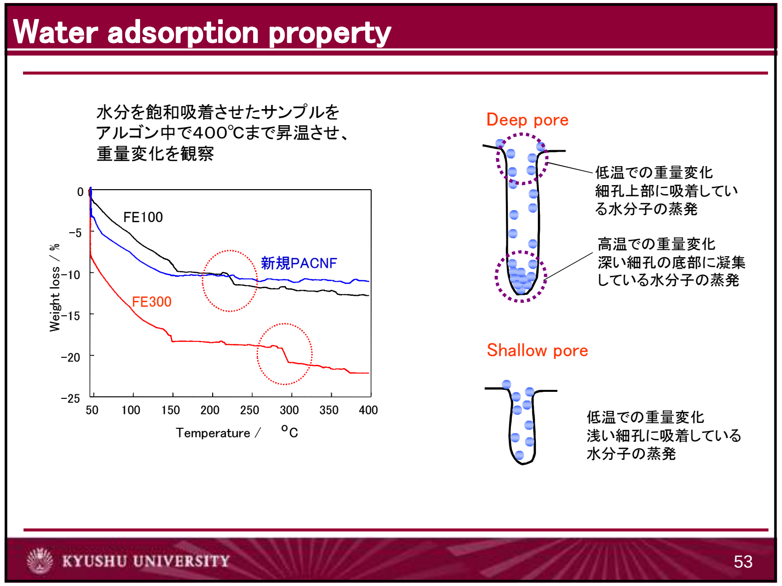
HCHO adsorption characteristics of PACNF in humidified atmosphere

RH	BET	Elemental analysis (wt%)						Microporous
	(m2 / g)	C	H	***N***	***Odiff***	ash	N / O	ratio (%)
90%	375	68.06	1.19	***18.02***	***11.41***	1.32	1.80	94.7%

Experimental
HCHO : 11 ppm
Sample weight : 0.05g
Gas flow rate : 100ml / ml
Humidity of condition : 0%

Experimental
HCHO : 11 ppm
Sample weight : 0.05g
Gas flow rate : 100ml / ml
Humidity of condition : 50%

**Under the circumstances of humidity (RH=50%),
PACNF shows specific prominent adsorption characteristics for formaldehyde.**



Water adsorption property

水分を飽和吸着させたサンプルをアルゴン中で400℃まで昇温させ、重量変化を観察

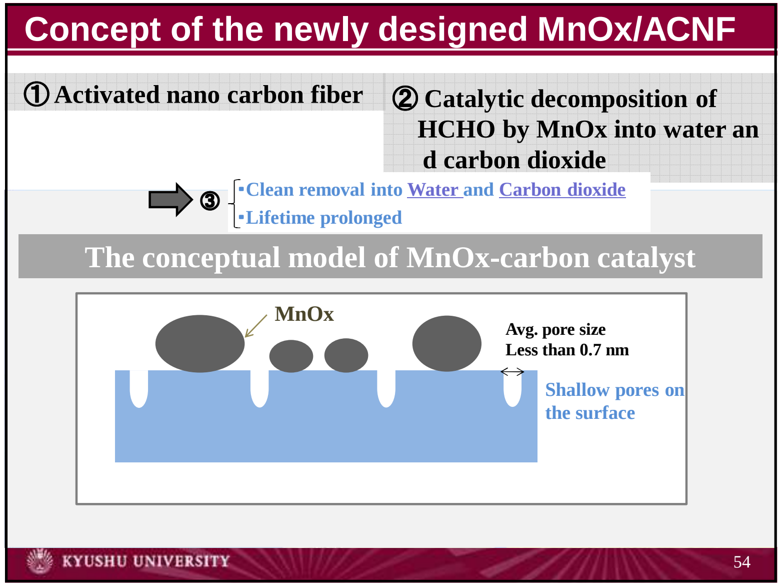
Deep pore

低温での重量変化
細孔上部に吸着している水分子の蒸発

高温での重量変化
深い細孔の底部に凝集している水分子の蒸発

Shallow pore

低温での重量変化
浅い細孔に吸着している水分子の蒸発



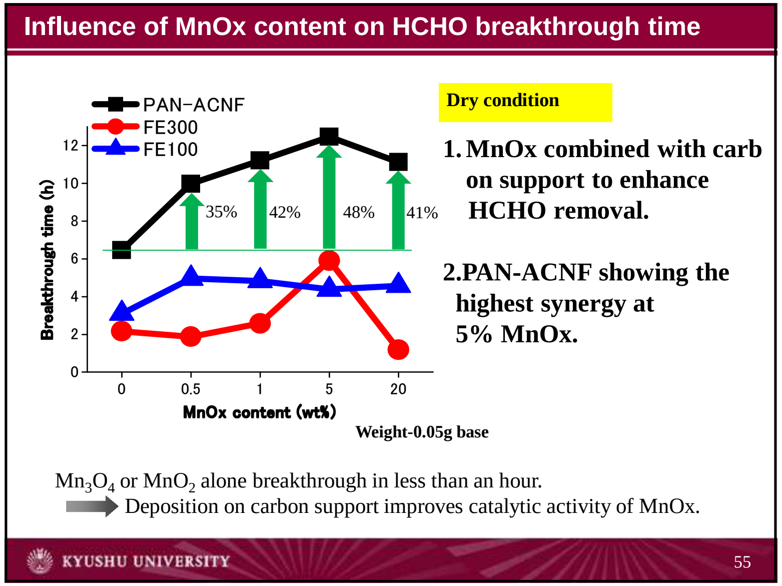
Concept of the newly designed MnOx/ACNF

① Activated nano carbon fiber

② Catalytic decomposition of HCHO by MnOx into water and carbon dioxide

➡ ③
- Clean removal into Water and Carbon dioxide
- Lifetime prolonged

The conceptual model of MnOx-carbon catalyst



Influence of MnOx content on HCHO breakthrough time

Dry condition

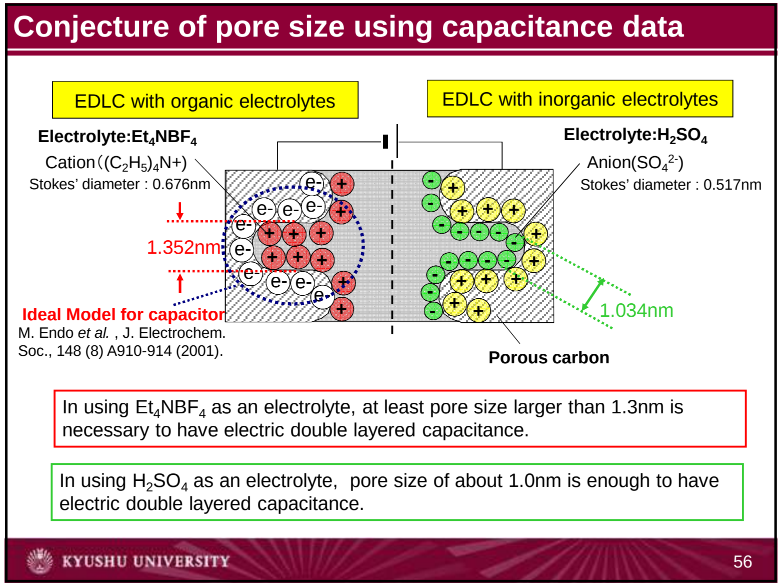
1. MnOx combined with carbon support to enhance HCHO removal.

2. PAN-ACNF showing the highest synergy at 5% MnOx.

Weight-0.05g base

Mn_3O_4 or MnO_2 alone breakthrough in less than an hour.

➡ Deposition on carbon support improves catalytic activity of MnOx.



Conjecture of pore size using capacitance data

EDLC with organic electrolytes

EDLC with inorganic electrolytes

Electrolyte:Et_4NBF_4

Cation（$(C_2H_5)_4N+$)
Stokes' diameter : 0.676nm

Electrolyte:H_2SO_4

Anion(SO_4^{2-})
Stokes' diameter : 0.517nm

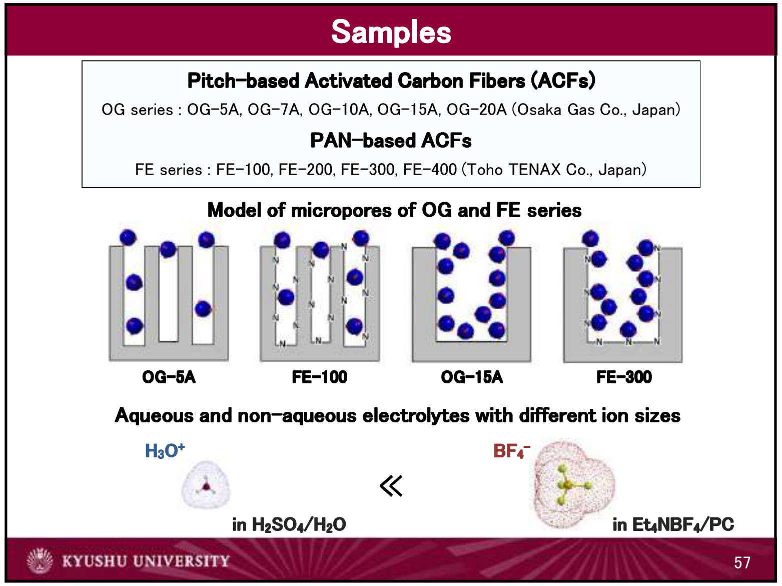
Ideal Model for capacitor

M. Endo *et al.* , J. Electrochem. Soc., 148 (8) A910-914 (2001).

Porous carbon

In using Et_4NBF_4 as an electrolyte, at least pore size larger than 1.3nm is necessary to have electric double layered capacitance.

In using H_2SO_4 as an electrolyte, pore size of about 1.0nm is enough to have electric double layered capacitance.



Samples

Pitch-based Activated Carbon Fibers (ACFs)

OG series : OG-5A, OG-7A, OG-10A, OG-15A, OG-20A (Osaka Gas Co., Japan)

PAN-based ACFs

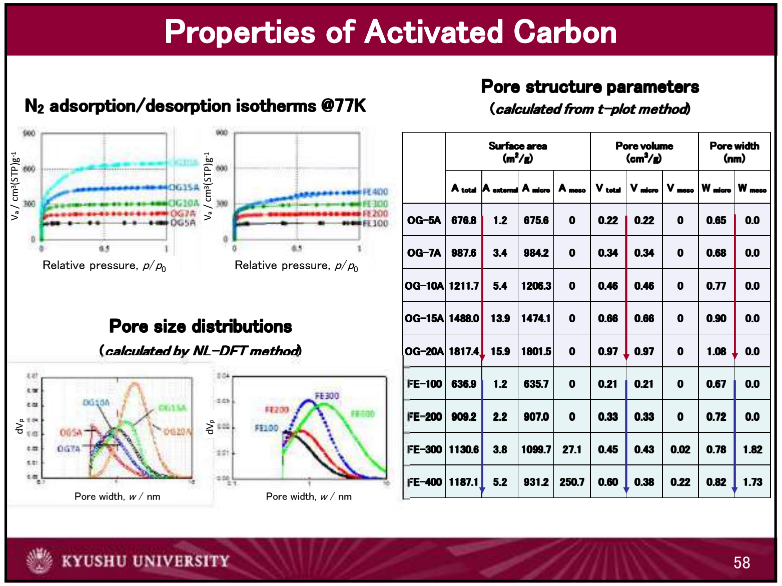
FE series : FE-100, FE-200, FE-300, FE-400 (Toho TENAX Co., Japan)

Model of micropores of OG and FE series

Aqueous and non-aqueous electrolytes with different ion sizes

in H_2SO_4/H_2O

in Et_4NBF_4/PC



Properties of Activated Carbon

N_2 adsorption/desorption isotherms @77K

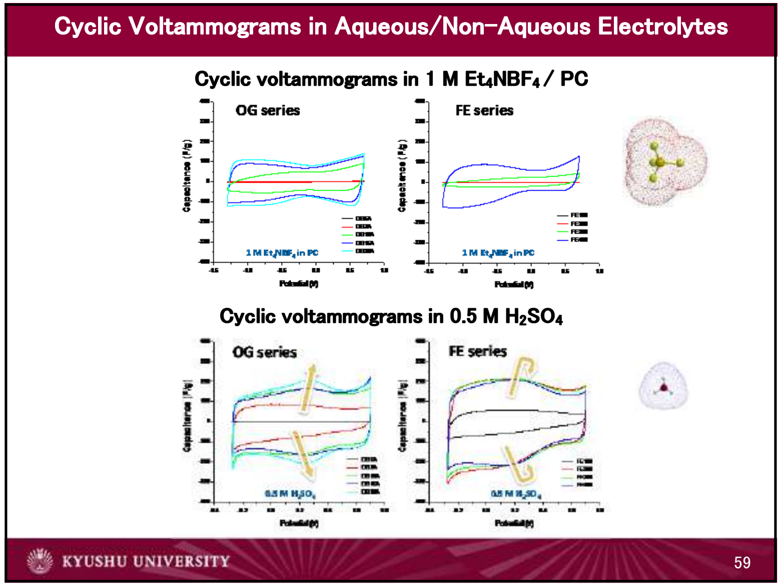
Pore size distributions

(*calculated by NL-DFT method*)

Pore structure parameters

(*calculated from t-plot method*)

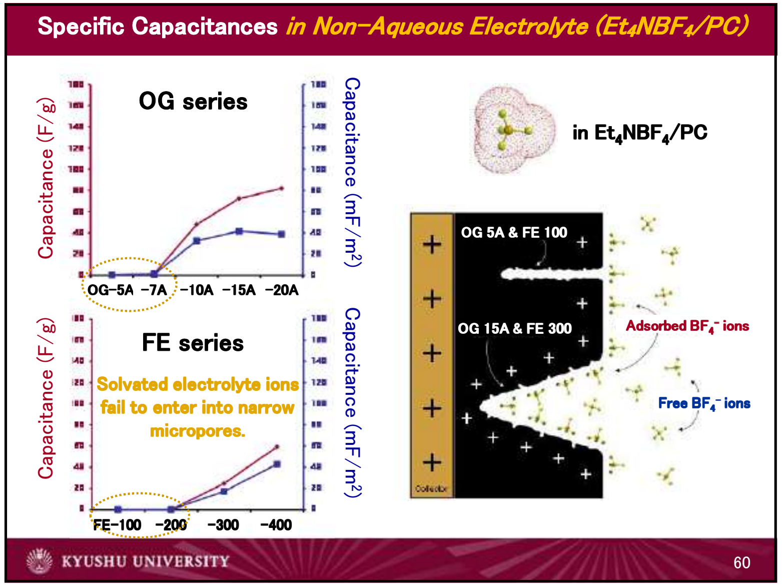
	Surface area (m^2/g)				Pore volume (cm^3/g)			Pore width (nm)	
	A_{total}	$A_{external}$	A_{micro}	A_{meso}	V_{total}	V_{micro}	V_{meso}	W_{micro}	W_{meso}
OG-5A	676.8	1.2	675.6	0	0.22	0.22	0	0.65	0.0
OG-7A	987.6	3.4	984.2	0	0.34	0.34	0	0.68	0.0
OG-10A	1211.7	5.4	1206.3	0	0.46	0.46	0	0.77	0.0
OG-15A	1488.0	13.9	1474.1	0	0.66	0.66	0	0.90	0.0
OG-20A	1817.4	15.9	1801.5	0	0.97	0.97	0	1.08	0.0
FE-100	636.9	1.2	635.7	0	0.21	0.21	0	0.67	0.0
FE-200	909.2	2.2	907.0	0	0.33	0.33	0	0.72	0.0
FE-300	1130.6	3.8	1099.7	27.1	0.45	0.43	0.02	0.78	1.82
FE-400	1187.1	5.2	931.2	250.7	0.60	0.38	0.22	0.82	1.73



Cyclic Voltammograms in Aqueous/Non-Aqueous Electrolytes

Cyclic voltammograms in 1 M Et_4NBF_4 / PC

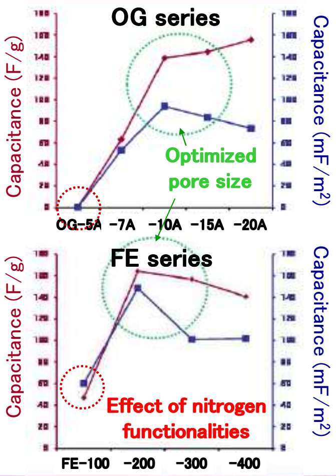
Cyclic voltammograms in 0.5 M H_2SO_4

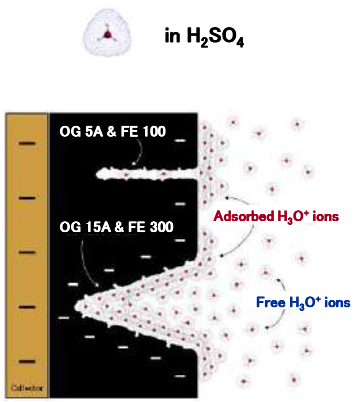


Specific Capacitances *in Non-Aqueous Electrolyte (Et_4NBF_4/PC)*

Solvated electrolyte ions fail to enter into narrow micropores.

Specific Capacitances *in Aqueous Electrolyte (H_2SO_4)*



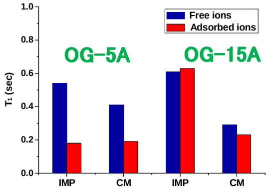


T_1 Values Measured from ^2H MAS Solid-State NMR Spectra

The shorter the T_1 value of relaxation time, the stronger the adsorption interaction between adsorbed electrolyte ions and carbon electrodes.

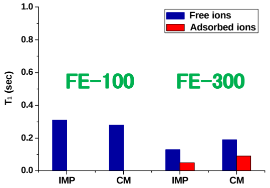
T_1 (sec) for OG series

OG-5A				OG-15A			
IMP		CM		IMP		CM	
Free	Adsorbed	Free	Adsorbed	Free	Adsorbed	Free	Adsorbed
0.54	0.18	0.41	0.19	0.61	0.63	0.29	0.23



T_1 (sec) for FE series

FE-100				FE-300			
IMP		CM		IMP		CM	
Free	Adsorbed	Free	Adsorbed	Free	Adsorbed	Free	Adsorbed
0.31	–	0.28	–	0.13	0.05	0.19	0.09



Steam vs. Chemical Activations

What is the difference:

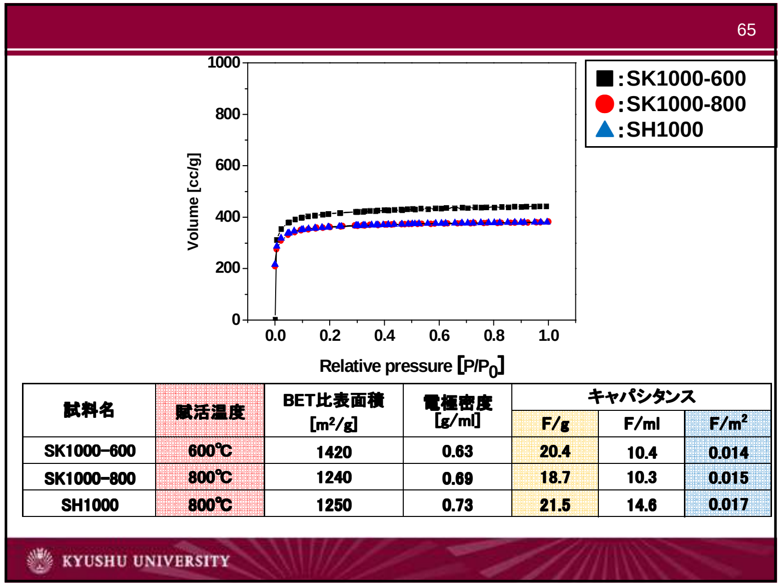
- Surface area, pore size and its distribution
- Surface compositions
- Surface structure (?)
- Cost
- Waste materials

Capacitance, cost, …

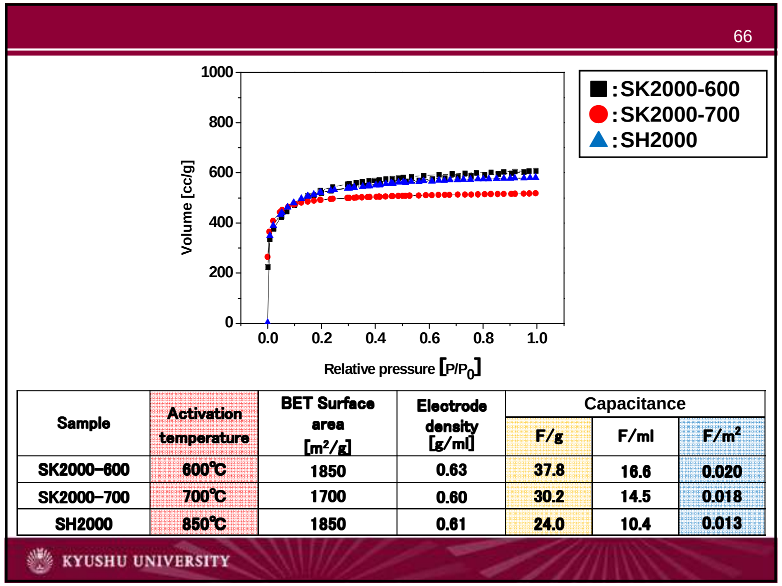
How to overcome the differences?

Structure of Activated Carbon

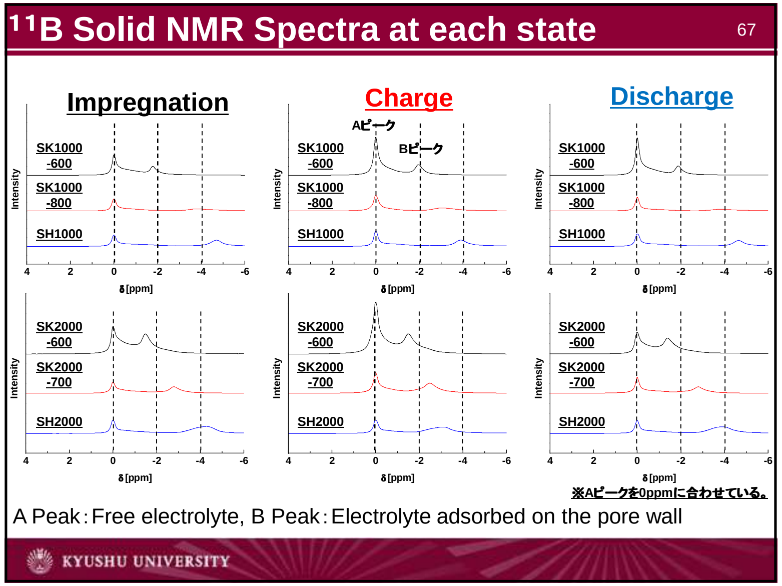
- **Surface Area, Pore: Depth & Volume**
  - Surface Structure
  - Surface Chemistry
  - Based and Edge Plane
  - Hetero atoms in Hexagon
- **Carbon Structure of Wall**
  - Nano, Micro, Macro Structure of Carbon Wall
  - -Graphitization Extent
  - -Domain Structure
    - → Density, Reactivity (Activated Surface)
    - Precursor : Structure and Reactivity



試料名	賦活温度	BET比表面積 [m²/g]	電極密度 [g/ml]	キャパシタンス		
				F/g	F/ml	F/m²
SK1000-600	600℃	1420	0.63	20.4	10.4	0.014
SK1000-800	800℃	1240	0.69	18.7	10.3	0.015
SH1000	800℃	1250	0.73	21.5	14.6	0.017

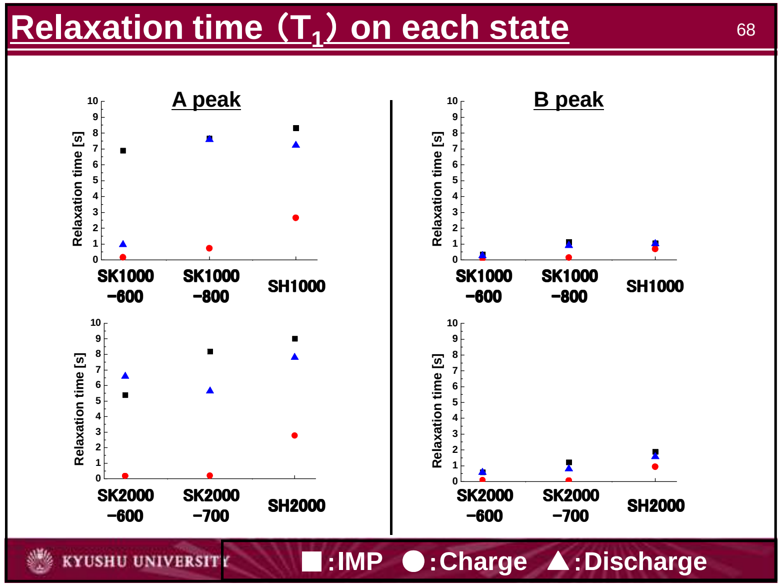


Sample	Activation temperature	BET Surface area [m²/g]	Electrode density [g/ml]	Capacitance		
				F/g	F/ml	F/m²
SK2000-600	600℃	1850	0.63	37.8	16.6	0.020
SK2000-700	700℃	1700	0.60	30.2	14.5	0.018
SH2000	850℃	1850	0.61	24.0	10.4	0.013

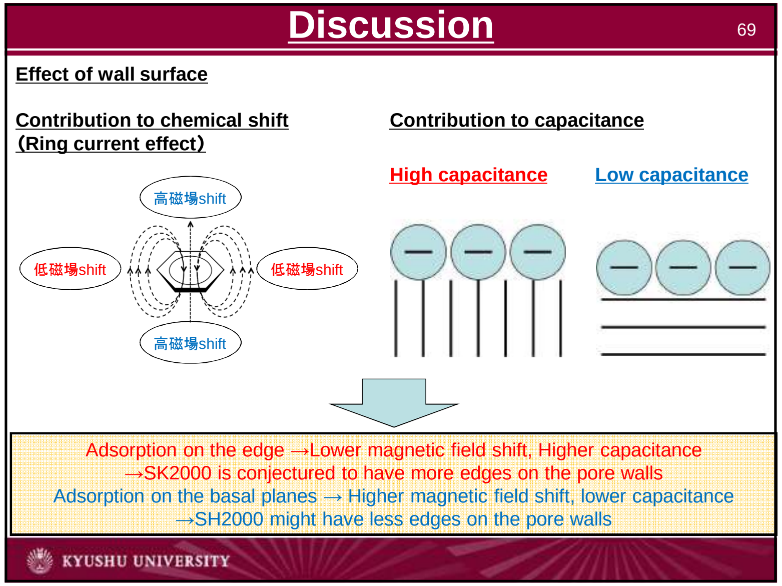

^{11}B Solid NMR Spectra at each state

※Aピークを0ppmに合わせている。

A Peak：Free electrolyte, B Peak：Electrolyte adsorbed on the pore wall


Relaxation time（T_1）on each state

■:IMP ●:Charge ▲:Discharge



Discussion

Effect of wall surface

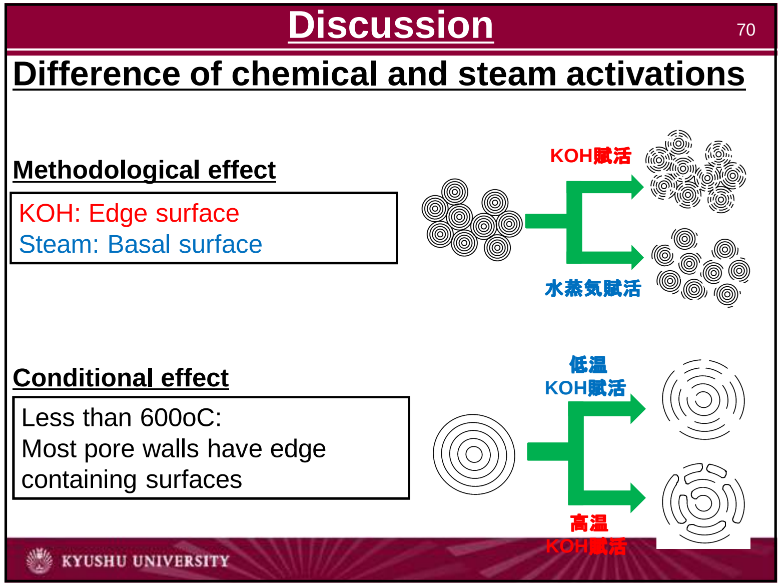
Contribution to chemical shift (Ring current effect)

Contribution to capacitance

High capacitance

Low capacitance

Adsorption on the edge →Lower magnetic field shift, Higher capacitance
→SK2000 is conjectured to have more edges on the pore walls
Adsorption on the basal planes → Higher magnetic field shift, lower capacitance
→SH2000 might have less edges on the pore walls



Discussion

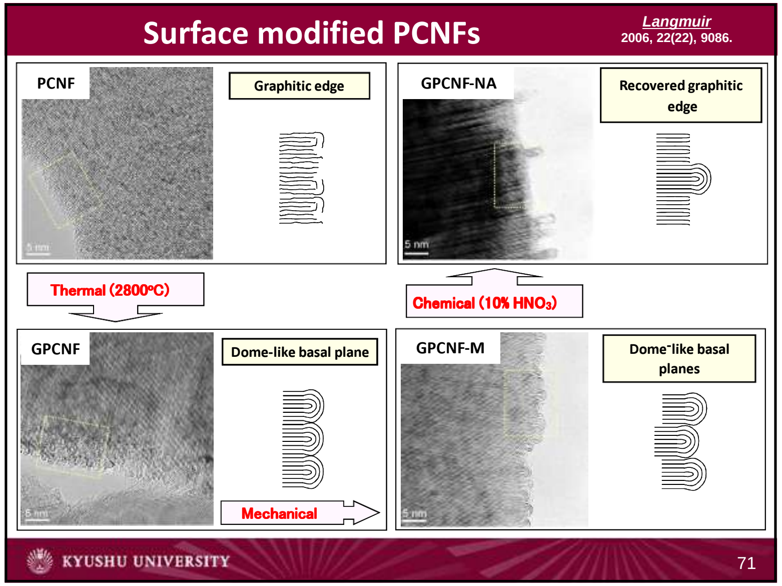
Difference of chemical and steam activations

Methodological effect

KOH: Edge surface
Steam: Basal surface

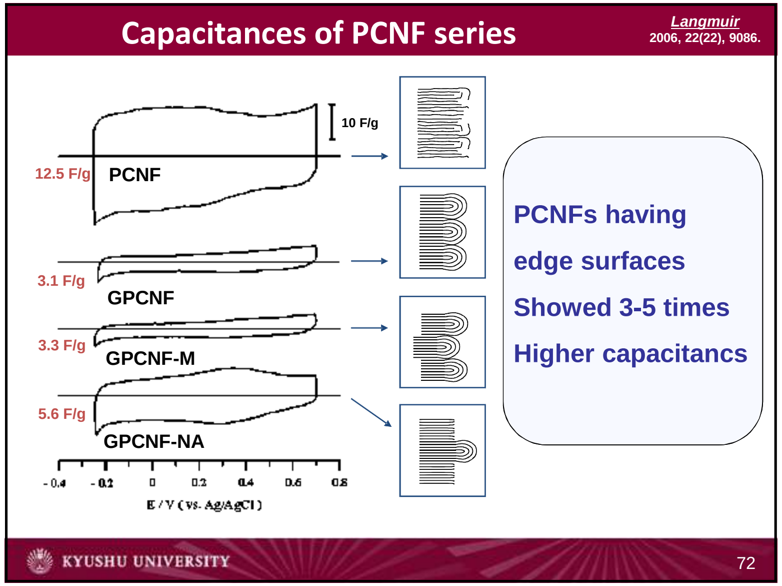
Conditional effect

Less than 600oC:
Most pore walls have edge containing surfaces



Surface modified PCNFs

Langmuir 2006, 22(22), 9086.



Capacitances of PCNF series

Langmuir 2006, 22(22), 9086.

PCNFs having edge surfaces Showed 3-5 times Higher capacitancs

Conclusion

- Carbon is a key material for energy and environmental devices.
- Full understanding of carbon structure is necessary for useful applications

- Korea has a lot of sources for carbon materials.
- No manpower and skill for carbon manufacturing.

University:

- ❖ Changing the consciousness
- ❖ Creation and leading of projects
- ❖ Manpower cultivation